Equilibrium

AP Chemistry · Vocabulary in context

Word	Example
reversible reaction /rɪ'vɜːsɪbl rɪ'ækʃn/	A **reversible reaction** can go both ways.
chemical equilibrium /'kemɪkl ˌiːkwɪ'lɪbrɪəm/	At **chemical equilibrium**, the amounts stop changing.
forward reaction /'fɔːwəd rɪ'ækʃn/	In the **forward reaction**, nitrogen and hydrogen make ammonia.
reverse reaction /rɪ'vɜːs rɪ'ækʃn/	In the **reverse reaction**, ammonia breaks down again.
concentration /ˌkɒnsən'treɪʃn/	The **concentration** of the acid is 2 $\mathrm{mol/dm^3}$. $2\ \mathrm{mol/dm^3}$
dynamic /daɪ'næmɪk/	Equilibrium is **dynamic**: both reactions keep going.
reactant /rɪ'æktənt/	A **reactant** is used up in a reaction.
product /'prɒdʌkt/	A **product** is made in a reaction.
reaction quotient /rɪ'ækʃn 'kwəʊʃənt/	The **reaction quotient** uses current concentrations to predict reaction direction.
equilibrium constant /ˌiːkwɪ'lɪbrɪəm 'kɒnstənt/	A large **equilibrium constant** means mostly products.
balanced equation	The **balanced equation** cannot tell you the orders.
partial pressure /'pɑːʃl 'preʃə/	The **partial pressure** of oxygen in air is about 21 kPa. 21 kPa

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molarity /mə'lærɪti/	**Molarity** is moles of solute per litre of solution.
coefficient /ˌkəʊɪ'fɪʃənt/	The number in front of the letter is the **coefficient**.
ICE table /aɪs 'teɪbl/	An **ICE table** shows initial, change and equilibrium amounts.
mole ratio /məʊl 'reɪʃɪəʊ/	Use the equation's **mole ratio** to calculate product amounts.
negligible /'neglɪdʒəbl/	For a heavy stone, air resistance is **negligible**.
particulate /pə'tɪkjʊleɪt/	A **particulate** diagram shows particles as small circles.
Le Chatelier's principle /lə 'tʃeɪtlɪəz 'prɪnsɪpl/	**Le Chatelier's principle** says equilibrium shifts to undo a change.
slightly soluble	Silver chloride is **slightly soluble**: only a tiny amount dissolves.
solubility product /ˌsɒljuː'bɪlɪti 'prɒdʌkt/	Silver chloride has a very small **solubility product**.
stoichiometry /ˌstəʊɪkɪ'ɒmətri/	**Stoichiometry** tells us how many moles of each substance react.
molar solubility /'məʊlə ˌsɒljuː'bɪlɪti/	**Molar solubility** is the moles that dissolve in one litre.
common-ion effect /'kɒmən 'aɪɒn ɪ'fekt/	The **common-ion effect** lowers solubility when the same ion is already present.
endothermic /ˌendəʊ'θɜːmɪk/	Melting ice is **endothermic**: it takes in heat.

Word	Example
catalyst /ˈkætəlɪst/	A **catalyst** speeds up a reaction without being used up.

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