Chemical Reactions

AP Chemistry · Vocabulary in context

Word	Example
chemical reaction /ˈkemɪkl rɪˈækʃn/	Digestion is a **chemical reaction**.
precipitate /prɪˈsɪpɪteɪt/	A white **precipitate** forms when silver nitrate meets chloride ions.
net ionic equation /net aɪˈɒnɪk ɪˈkweɪʒn/	A **net ionic equation** leaves out spectator ions.
spectator ions /spekˈteɪtəˈaɪɒnz/	Na+ and Cl− are **spectator ions**: they do not change. Na+ Cl−
complete ionic equation	A **complete ionic equation** shows every ion in solution.
physical change /ˈfɪzɪkl tʃeɪndʒ/	Melting ice is a **physical change**: no new substance forms.
chemical change /ˈkemɪkl tʃeɪndʒ/	Burning wood is a **chemical change**: new substances form.
Stoichiometry /ˌstəʊɪkɪˈɒmətri/	**Stoichiometry** uses mole ratios to find amounts.
limiting reactant /ˈlɪmɪtɪŋ rɪˈæktənt/	The **limiting reactant** runs out first.
theoretical yield /θɪəˈretɪkl jiːld/	The **theoretical yield** is the maximum product amount predicted by calculation.
percent yield	The **percent yield** compares the real amount with the expected amount.
titration /taɪˈtreɪʃn/	A **titration** finds the exact volume of acid needed.

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equivalence point /ɪˈkwɪvələns pɔɪnt/	At the **equivalence point**, the acid and base have exactly reacted.
titrant	The **titrant** is the solution in the burette.
analyte	The **analyte** is the solution in the flask.
combustion /kəmˈbʌstʃn/	**Combustion** of petrol powers a car.
acid /ˈæsɪd/	Lemon juice is an **acid**.
base /beɪs/	Sodium hydroxide is a **base**.
conjugate acid–base pairs	HCl and Cl^- are **conjugate acid–base pairs**. HCl Cl^-
redox /rɪˈdɒks/	A **redox** reaction has oxidation and reduction together.
Oxidation /ˌɒksɪˈdeɪʃn/	**Oxidation** is the loss of electrons.
reduction /rɪˈdʌkʃn/	**Reduction** is the gain of electrons.
half-reactions	Split a redox reaction into two **half-reactions**.