30. Hydrocarbons

Starter Quiz · 课前小测

A-Level Chemistry Name: _____ Score: _____

1. Benzene reacts mostly by electrophilic substitution because:
   - ☐ this keeps the stable delocalised ring, while addition would destroy it
   - ☐ the ring is unstable
   - ☐ it has no electrons
   - ☐ addition is impossible
2. A halogen reacts in the ring (not the side-chain) when there is:
   - ☐ a halogen-carrier catalyst ($AlCl_3$) and no UV light
   - ☐ UV light and no catalyst
   - ☐ only heat
   - ☐ water present
3. Nitration of benzene uses:
   - ☐ concentrated HNO_3 and concentrated H_2SO_4
   - ☐ bromine water
   - ☐ dilute NaOH
   - ☐ hydrogen and a catalyst
4. Under UV light with no catalyst, the halogen substitutes:
   - ☐ in the side-chain (free-radical substitution)
   - ☐ in the ring
   - ☐ nowhere
   - ☐ at the catalyst
5. A Friedel–Crafts reaction (with $AlCl_3$) is used to:
   - ☐ add an alkyl or acyl group to the ring
   - ☐ remove a hydrogen as a radical
   - ☐ oxidise the ring
   - ☐ hydrogenate the ring
6. Groups already on a benzene ring direct where the next group attaches.

   ☐ True ☐ False

7. Benzene reacts mainly by electrophilic substitution, not addition, because that keeps its stable delocalised ring intact.

   ☐ True ☐ False
8. An –OH group already on the ring directs the next substituent to:
   - ☐ positions 2 and 4
   - ☐ position 3 only
   - ☐ no particular position
   - ☐ the side-chain
9. In electrophilic substitution, after the ring bonds to the electrophile:
   - ☐ an H^+ is lost, restoring the stable ring
   - ☐ the ring opens up
   - ☐ another electrophile adds
   - ☐ a radical forms
10. Arenes such as benzene typically react by:
    - ☐ electrophilic substitution
    - ☐ electrophilic addition
    - ☐ nucleophilic substitution
    - ☐ free-radical addition

A-Level Chemistry

1. **this keeps the stable delocalised ring, while addition would destroy it**

   Substitution preserves the aromatic stabilisation; addition would break up the delocalised system.

2. **a halogen-carrier catalyst ($AlCl_3$) and no UV light**

   The catalyst makes the electrophile for ring substitution; UV with no catalyst gives side-chain free-radical substitution.

3. **concentrated HNO_3 and concentrated H_2SO_4**

   The conc. acid mixture generates the NO_2^+ electrophile, giving nitrobenzene.

4. **in the side-chain (free-radical substitution)**

   UV light starts free-radical substitution in the alkyl side-chain.

5. **add an alkyl or acyl group to the ring**

   Friedel–Crafts alkylation/acylation adds a carbon-based group to benzene.

6. **True**

   This is the directing effect.

7. **True**

   Addition would break the delocalised ring and lose its stability, so substitution (swapping an H) is favoured.

8. **positions 2 and 4**

   $–NH_2$, –OH and –R are 2,4-directing; $–NO_2$, –COOH and –COR are 3-directing.

9. **an H^+ is lost, restoring the stable ring**

   Losing H^+ from that carbon restores the delocalised ring, completing the substitution.

10. **electrophilic substitution**

    Substitution keeps the stable ring intact.