

Multiple Choice Answers

Q18: D

Multiple Choice Answers

Q18: D

Q23: B

Multiple Choice Answers

Q16: A

2(a)(i)	M1 nuclear charge increases (across period) AND similar shielding M2 greater / stronger attraction to nucleus														
2(a)(ii)	M1 Al loses (3) electrons / Al is 3+, and P gains (3) electrons / P is 3– M2 Al (3+) has (one) fewer shell (than P(3–))														
2(b)(i)	<table><tr><td>M1</td><td>(+)1</td><td>(+)2</td><td>(+)3</td><td>(+)4</td><td>(+)5</td><td>(+)6</td></tr><tr><td>M2</td><td>12-14</td><td>8-10</td><td>—</td><td>—</td><td>0-4</td><td>0-4</td></tr></table>	M1	(+)1	(+)2	(+)3	(+)4	(+)5	(+)6	M2	12-14	8-10	—	—	0-4	0-4
M1	(+)1	(+)2	(+)3	(+)4	(+)5	(+)6									
M2	12-14	8-10	—	—	0-4	0-4									
2(b)(ii)	they are insoluble														
2(c)(i)	$\text{Na}_2\text{O} + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{H}_2\text{O}$														
2(c)(ii)	$\text{Al}_2\text{O}_3 + 2\text{NaOH} \rightarrow 2\text{NaAlO}_2 + \text{H}_2\text{O}$														
2(d)(i)	(thermal stability) increases (down the group)														
2(d)(ii)	NO_2 AND O_2														

2(a)(i)	delocalised electrons move (through the structure / between layers)
2(a)(ii)	P / S / Cl have no charge carriers OR (no ions and) no delocalised electrons OR no mobile / moving ions
2(b)	M1 Si has (lots of) strong covalent bonds broken M2 P / S / Cl (only) has weak Van der Waals' (forces) broken

2(c)	<table><tr><td>M1</td><td>+1</td><td>+2</td><td>+3</td><td>+4</td><td>+5</td></tr><tr><td>M2</td><td>7</td><td>6.5</td><td>0–4</td><td>0–4</td><td>0–4</td></tr><tr><td>M3</td><td>I</td><td>I</td><td>I</td><td>C</td><td>C</td></tr><tr><td>M4</td><td>G</td><td>G</td><td>G</td><td>S</td><td>S</td></tr></table>	M1	+1	+2	+3	+4	+5	M2	7	6.5	0–4	0–4	0–4	M3	I	I	I	C	C	M4	G	G	G	S	S
M1	+1	+2	+3	+4	+5																				
M2	7	6.5	0–4	0–4	0–4																				
M3	I	I	I	C	C																				
M4	G	G	G	S	S																				
2(d)(i)	$2\text{Al} + 3\text{Cl}_2 \rightarrow 2\text{AlCl}_3$																								
2(d)(ii)	$\text{PCl}_5 + 4\text{H}_2\text{O} \rightarrow \text{H}_3\text{PO}_4 + 5\text{HCl}$																								
2(e)(i)	rate of forward and backward (reaction) are the same / equal OR (ratio of) concentrations of reactants and products are constant																								
2(e)(ii)	closed system																								
2(e)(iii)	M1 (mixture / solution becomes more) yellow M2 (forward) reaction is exothermic (equilibrium shifts left) OR equilibrium / reaction shifts (left) to absorb (additional) heat / energy M3 (mixture / solution becomes more) red M4 fewer moles of gas on right-hand side OR (gas on left and) no gas on the right																								
2(f)(i)	$\text{MgCl}_2(\text{aq}) + \text{Na}_2\text{CO}_3(\text{aq}) \rightarrow \text{MgCO}_3(\text{s}) + 2\text{NaCl}(\text{aq})$ M1 formulae in balanced equation M2 state symbols																								
2(f)(ii)	MgO / magnesium oxide AND CO ₂ / carbon dioxide																								
2(f)(iii)	(thermal stability) increases (down the group)																								

2(a)	M1 Al_2O_3 is ionic M2 P_4O_{10} is a (simple) molecule M3 ionic bonds are stronger than intermolecular forces
2(b)(i)	$\text{mol}^{-2} \text{ dm}^6$
2(b)(ii)	M1 at equilibrium: moles of CO = 0.120 AND moles of H_2 = 0.240 M2 $[\text{CH}_3\text{OH}] = 0.056$ AND $[\text{CO}] = 0.024$ AND $[\text{H}_2] = 0.048$ M3 use of K_c with correct evaluation = 1.01×10^3
2(b)(iii)	no effect AND K_c is not affected by pressure
2(c)(i)	increases rate by providing a different mechanism with a lower activation energy
2(c)(ii)	$2\text{CH}_3\text{OH} \rightarrow \text{CH}_3\text{OCH}_3 + \text{H}_2\text{O}$
2(c)(iii)	$\text{P}_4\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$