

Multiple Choice Answers

Q30: D

Q31: A

Multiple Choice Answers

Q30: A

Q31: C

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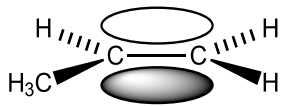
Q31: D

Q32: A

Q38: B

4(a)(i)	free-radical substitution
4(a)(ii)	M1 $\text{C}_3\text{H}_8 + \text{Cl}\cdot \rightarrow \text{CH}_3\text{CH}_2\text{CH}_2\cdot + \text{HCl}$ OR $\text{C}_3\text{H}_8 + \text{Cl}\cdot \rightarrow \text{C}_3\text{H}_7\cdot + \text{HCl}$ M2 $\text{CH}_3\text{CH}_2\text{CH}_2\cdot + \text{Cl}_2 \rightarrow \text{CH}_3\text{CH}_2\text{CH}_2\text{Cl} + \text{Cl}\cdot$ OR $\text{C}_3\text{H}_7\cdot + \text{Cl}_2 \rightarrow \text{C}_3\text{H}_7\text{Cl} + \text{Cl}\cdot$ etc M3 $\text{CH}_3\text{CH}_2\text{CH}_2\cdot + \text{Cl}\cdot \rightarrow \text{CH}_3\text{CH}_2\text{CH}_2\text{Cl}$ OR $\text{C}_3\text{H}_7\cdot + \text{Cl}\cdot \rightarrow \text{CH}_3\text{CH}_2\text{CH}_2\text{Cl}$
4(a)(iii)	homolytic
4(a)(iv)	1,2-dichloropropane
4(a)(v)	M1 (type of structural isomerism) positional M2 (type of stereoisomerism) optical
4(b)	sodium hydroxide AND water
4(c)(i)	orange to green
4(c)(ii)	$\text{CH}_3\text{CH}_2\text{CH}_2\text{OH} + 2[\text{O}] \rightarrow \text{CH}_3\text{CH}_2\text{COOH} + \text{H}_2\text{O}$
4(d)(i)	condensation
4(d)(ii)	M1 absorption labelled at approx. 1750 cm^{-1} identified as C=O M2 absorption labelled at approx. 1650 cm^{-1} identified as C=C M3 absorption labelled at approx. 1170 cm^{-1} identified as C—O
4(d)(iii)	$(0.31 / 4.7 \times 100 / 1.1) = 6$ carbon atoms

3(a)(i)	
3(a)(ii)	yellow precipitate
3(a)(iii)	elimination

3(a)(iv)	
3(b)(i)	$(\text{CH}_3)_2\text{CHOH} + \text{Na} \rightarrow (\text{CH}_3)_2\text{CHO}^{(-)}\text{Na}^{(+)} + \frac{1}{2} \text{H}_2$
3(b)(ii)	M1 correct dipole on $\delta^+\text{C}-\text{Br}\delta^-$ AND curly arrow from C—Br bond to Br M2 intermediate  M3 curly arrow from lone pair of $(\text{CH}_3)_2\text{CHO}^-$ to C^+
3(b)(iii)	M1 rate would decrease M2 C—Br is weaker (than C—Cl, so breaks more easily)
3(b)(iv)	$\text{C}_6\text{H}_{14}\text{O} + 9\text{O}_2 \rightarrow 6\text{CO}_2 + 7\text{H}_2\text{O}$