2(a)	the (3)d and (4)s sub-shells / orbitals / electrons are close / similar in energy
2(b)	species or ion formed by a central metal atom/ion **AND** surrounded by/bonded to (one or more) ligands
2(c)(i)	NaOH / OH^-(aq) **AND** precipitation / ligand exchange / deprotonation / acid–base
2(c)(ii)	**M1**: $Fe(H_2O)_4(OH)_2$ is oxidised (from + 2 to + 3 on standing) **M2**: $E^\ominus$ of $Fe(H_2O)_3(OH)_3$ is more negative than $E^\ominus$ of O_2 **OR** $E^\ominus_{cell}$ = 0.40 – (–0.56) = (+)0.96 V **M3**: $4Fe(H_2O)_4(OH)_2 + O_2 \rightarrow 4Fe(H_2O)_3(OH)_3 + 2H_2O$
2(d)	$\Delta G^\ominus = -2 \times 1.67 \times 96\,500 = -322\,310$ J mol^{-1} **M1:** $\Delta G^\ominus = -nE^\ominus_{cell}F$ **AND** n = 2 **OR** $-2 \times 1.67 \times 96\,500$ **M2:** $\Delta G^\ominus$ = –322.3 kJ mol^{-1}