

6(a)(i)	$\text{Ni}(\text{CO})_4$
6(a)(ii)	- ethanedioate given as C_2O_4 - formula shows one Ni and three ligands - charge = 4^- any two [1], all three [2] $\text{Ni}(\text{C}_2\text{O}_4)_3^{4-}$ scores [2]
6(a)(iii)	F AND optical
6(b)(i)	N atom can donate one lone pair of electrons
6(b)(ii)	$[\text{Cd}(\text{en})_2]^{2+}$ AND has a larger K_{stab} / is more stable
6(b)(iii)	$\frac{[\text{Cd}(\text{CH}_3\text{NH}_2)_4]^{2+}}{[\text{Cd}^{2+}][\text{CH}_3\text{NH}_2]^4}$