# 26 Rate equations and orders – Part 1 —Answers

## Answers

Work through these after trying the questions yourself.

**26.1** How the rate depends on the reactant concentrations.
**26.2** The order of reaction with respect to A.
**26.3** The sum of the individual orders $(m + n)$.
**26.4** By experiment.
**26.5** Zero order.

**26.6** The order with respect to B is 2. Since rate $\propto [B]^2$, doubling [B] multiplies the rate by $2^2 = 4$ —the rate becomes four times larger.