

23 Born-Haber cycles – Part 2 —Answers

Answers

Work through these after trying the questions yourself.

23.6 Any one unknown enthalpy change (such as the lattice energy).

23.7 Atomisation or ionisation.

23.8 The lattice energy.

23.9 Hess's law.

23.10 Stronger (more negative).

23.11 Lithium and fluoride ions are much *smaller* than potassium and iodide ions. Smaller ions sit closer together in the lattice, so the electrostatic attraction between them is stronger and more energy is released —a more negative lattice energy.