

6 Oxidation number and redox – Part 2

—Answers

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Work through these after trying the questions yourself.

6.1 They must be equal.

6.2 (a) by 1 \ (b) oxidised (the number rises).

6.3 Five iron atoms (rise of 1 each $\times 5 = 5$, matching the fall of 5).

6.4 A redox reaction in which the same element is both oxidised and reduced at the same time.

6.5 (a) 0 \ (b) -1 \ (c) $+1$.

6.6 Chlorine starts at 0 in Cl_2 . In NaCl it becomes -1 (reduced) and in NaClO it becomes $+1$ (oxidised). The *same* element is both oxidised and reduced, so it is disproportionation.