

Solutions

Each answer shows the steps, then the **result**.

2.1

- the electrostatic attraction between positive metal ions and delocalised (sea of) electrons

2.2

- delocalised electrons

3.1 B

- Delocalised electrons are free to move and carry charge

3.2

- the delocalised electrons are free to move through the metal; they carry charge when a voltage is applied

3.3

- magnesium ions have a 2+ charge (vs 1+ for sodium) and contribute more delocalised electrons; the stronger attraction needs more energy to break, so a higher melting point