

3.1 Electronegativity and bonding

Name: _____ Class: _____ Date: _____

Total: 9 marks

KEY WORDS electronegativity 电负性 • electron 电子 • ion 离子 • shielding 屏蔽
• atomic radius 原子半径

Objective

Build the skills to answer exam questions on **electronegativity and bonding** — defining **electronegativity** 电负性, its trends, and using **Pauling** 鲍林 values to predict bond type.

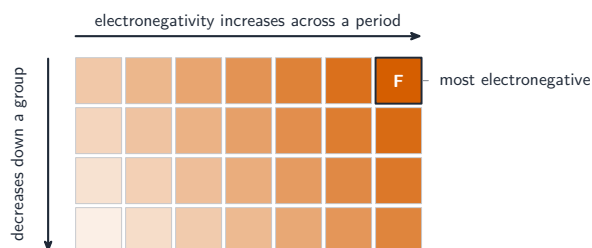
You must be able to:

- define electronegativity and explain the factors that affect it
- state and explain the trends across a period and down a group
- use electronegativity differences to predict ionic or covalent bonding

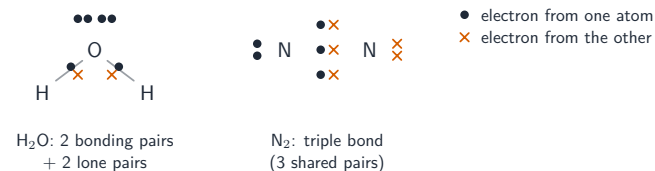
1 Worked examples

■ Electronegativity and its trends

Electronegativity is the power of an atom to attract the **electrons** in a bond towards itself. It depends on **nuclear charge** (more → higher), **atomic radius** (smaller → higher) and **shielding** (more → lower).



Electronegativity rises across a period and falls down a group — fluorine is the most electronegative



Electronegativity sets where a bond sits on the ionic-covalent scale

■ Predicting bond type

A **large** difference in Pauling values → **ionic**; a **small** difference → **covalent**.

2 Practice

2.1 Define **electronegativity**.

[1]

2.2 State the most electronegative element.

[1]

3 Exam-style questions

3.1 Electronegativity increases across Period 3 because:

[1]

A	B
C	D

- A the atomic radius increases
- B the nuclear charge increases while shielding stays similar
- C more shells are added
- D the atoms gain electrons

3.2 Two elements have a **large** difference in electronegativity. The bond between them is likely to be:

[1]

A	B
C	D

- A metallic
- B ionic
- C pure covalent
- D a hydrogen bond

3.3 Explain why electronegativity decreases down Group 17. [3]

3.4 Using electronegativity, explain why the bond in HCl is polar. [2]
