

12.1 Nitrogen and sulfur

Name: _____ Class: _____ Date: _____

Total: 12 marks

KEY WORDS acid rain 酸雨 • bond energy 键能 • triple bond 三键 • basicity 碱性
• coordinate bond 配位键

Objective

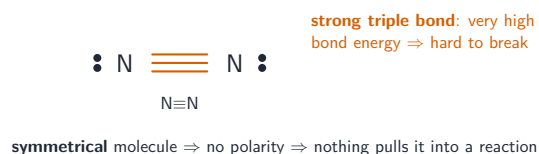
Build the skills to answer exam questions on **nitrogen and sulfur** — the unreactivity of nitrogen, the **basicity** 碱性 of **ammonia** 氨, oxides of nitrogen, and **acid rain** 酸雨.

You must be able to:

- explain the lack of reactivity of nitrogen
- explain the basicity of ammonia and the formation of the ammonium ion
- describe the sources, catalytic removal, and role in acid rain of nitrogen oxides

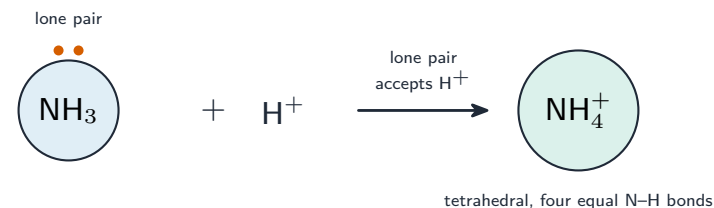
1 Worked examples

■ Unreactive nitrogen



N_2 is unreactive: a very strong triple bond (high bond energy) and a symmetrical, non-polar molecule

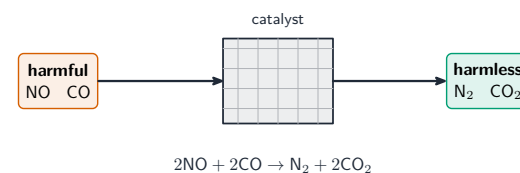
■ Ammonia as a base



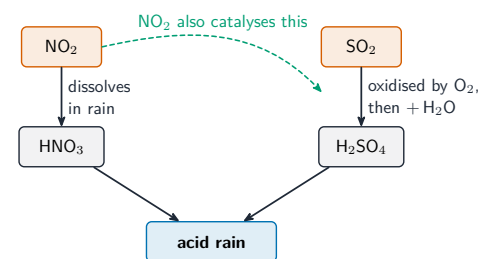
Ammonia is a base because the lone pair on N accepts a proton: $\text{NH}_3 + \text{H}^+ \rightarrow \text{NH}_4^+$

The new N–H bond is a **coordinate bond** (both electrons from N). Warming an ammonium salt with a base displaces ammonia: $\text{NH}_4\text{Cl} + \text{NaOH} \rightarrow \text{NaCl} + \text{NH}_3 + \text{H}_2\text{O}$.

■ Nitrogen oxides and acid rain



A catalytic converter: $2\text{NO} + 2\text{CO} \rightarrow \text{N}_2 + 2\text{CO}_2$



NO_2 forms nitric acid directly and catalyses the oxidation of SO_2 to sulfuric acid — both cause acid rain

2 Practice

2.1 State **two** reasons why nitrogen gas is unreactive. [2]

2.2 Write the equation for ammonia acting as a base towards H^+ . [1]

3 Exam-style questions

3.1 Ammonia acts as a Brønsted–Lowry base because: [1]

- | | |
|---|---|
| A | B |
| C | D |
- A it donates a proton
 - B the lone pair on nitrogen accepts a proton
 - C it is a gas
 - D it contains hydrogen

3.2 In a catalytic converter, nitrogen monoxide is removed by reacting with: [1]

- | | |
|---|---|
| A | B |
| C | D |
- A oxygen
 - B carbon monoxide
 - C water
 - D sulfur dioxide

3.3 (a) Describe how nitrogen monoxide is formed inside a car engine, and write an equation. [2]

(b) Explain the **two** ways that oxides of nitrogen contribute to acid rain. [3]

3.4 Describe a chemical test for an ammonium salt. [2]
