

34.4 Amino acids

Name: _____ Class: _____ Date: _____

Total: 10 marks

Objective

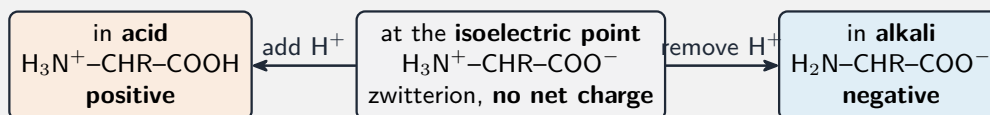
Build the skills to answer exam questions on **amino acids** 氨基酸—**zwitterions** 两性离子, the **isoelectric point** 等电点, **peptide bonds** 肽键 and **electrophoresis** 电泳.

You must be able to:

- explain how an amino acid forms a zwitterion and how its charge changes with pH
- describe peptide-bond formation by condensation
- explain how electrophoresis separates amino acids

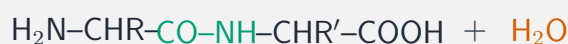
1 Worked examples

■ Charge with pH; peptide bonds



Charge depends on pH: positive in acid, the neutral zwitterion at the isoelectric point, negative in alkali

- in **acid** (low pH): $-\text{COO}^-$ gains H^+ → ion is **positive**.
- in **alkali** (high pH): $-\text{NH}_3^+$ loses H^+ → ion is **negative**.
- at the **isoelectric point**: mostly the neutral zwitterion, **no net charge**.



the **peptide bond** (an amide link)

Two amino acids condense (losing water) to form a peptide (amide) bond

2 Practice

2.1 State what is meant by a **zwitterion**. [1]

2.2 Name the type of bond formed when two amino acids join. [1]

3 Exam-style questions

3.1 At a pH **below** its isoelectric point, an amino acid carries an overall: [1]

- **A** positive charge
 - **B** negative charge
 - **C** zero charge
 - **D** double negative charge
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3.2 In electrophoresis, an amino acid that is **negatively** charged moves towards the: [1]

- **A** negative electrode
 - **B** positive electrode
 - **C** neither electrode
 - **D** both electrodes
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3.3 Explain, using equations or structures, how the charge on the amino acid glycine ($\text{H}_2\text{NCH}_2\text{COOH}$) changes between low pH and high pH. [3]

3.4 Two amino acids are mixed and held at a pH equal to the isoelectric point of one of them. Explain what happens to each in electrophoresis. [3]

4 Go further

You are now ready for the real exam questions on this subtopic. Open the **34.4 Amino acids** past-paper sheet in the Library, or try these in **Practice mode**:

- 9701/42 N25 —Q8 (amino acids)

Solutions

2.1 an ion with both a positive ($-\text{NH}_3^+$) and a negative ($-\text{COO}^-$) charge but no overall charge.

2.2 a peptide bond (an amide link).

3.1 A. Below the isoelectric point, extra H^+ gives an overall positive charge.

3.2 B. A negative ion is attracted to the positive electrode.

3.3 at low pH the $-\text{COO}^-$ is protonated to $-\text{COOH}$, giving $\text{H}_3\text{N}^+\text{CH}_2\text{COOH}$, overall positive; at high pH the $-\text{NH}_3^+$ loses H^+ to give $\text{H}_2\text{NCH}_2\text{COO}^-$, overall negative; between them lies the zwitterion $\text{H}_3\text{N}^+\text{CH}_2\text{COO}^-$ at the isoelectric point.

3.4 the amino acid at its isoelectric point is a neutral zwitterion, so it does not move; the other amino acid is charged at that pH; so it migrates towards the oppositely charged electrode, separating the two.