

18.1 Carboxylic acids

Name: _____ Class: _____ Date: _____

Total: 11 marks

KEY WORDS esterification 酯化 • neutralisation 中和 • reflux 回流 • nitrile 腈 • alcohol 醇

Objective

Build the skills to answer exam questions on **carboxylic acids** 羧酸 — making them, and their reactions as **acids** and to change the functional group.

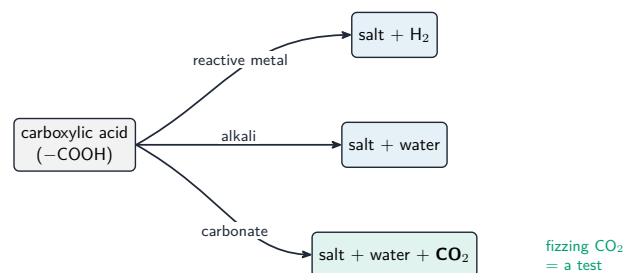
You must be able to:

- recall how carboxylic acids are made (oxidation, hydrolysis of nitriles/esters)
- describe the reactions with metals, alkalis and carbonates
- describe esterification and reduction by LiAlH_4

1 Worked examples

■ Making and reacting carboxylic acids

A **carboxylic acid** has the $-\text{COOH}$ group (a weak acid). Made by: oxidising a primary alcohol/aldehyde (reflux); hydrolysing a nitrile or ester then acidifying.



Acid reactions: with a metal $\rightarrow$ salt + H_2 ; with an alkali $\rightarrow$ salt + water; with a carbonate $\rightarrow$ salt + water + CO_2 (the fizz is a test)



the $-\text{OH}$ (acid) and $-\text{H}$ (alcohol) leave as **water**; the $\text{C}-\text{O}-\text{C}$ ester link forms
(concentrated H_2SO_4 catalyst; reversible)

Carboxylic acids form esters with alcohols

Two reactions change the functional group: **esterification** with an alcohol (conc. H_2SO_4) $\rightarrow$ ester; **reduction** by $\text{LiAlH}_4 \rightarrow$ primary alcohol.

2 Practice

2.1 Write the equation for ethanoic acid reacting with sodium carbonate. [1]

2.2 State a chemical test, with result, that shows a compound is a carboxylic acid. [2]

3 Exam-style questions

3.1 Which reagent reduces a carboxylic acid to a primary alcohol? [1]

- A NaBH_4
- B LiAlH_4
- C acidified $\text{K}_2\text{Cr}_2\text{O}_7$
- D concentrated H_2SO_4

3.2 Ethanoic acid reacts with magnesium to give: [1]

- A an ester and water
- B a salt and hydrogen
- C a salt, water and CO_2
- D no reaction

3.3 (a) Write the equation for the reaction of ethanoic acid with sodium hydroxide and name the type of reaction. [2]

(b) State the reagents and conditions to make ethanoic acid from ethanol. [2]

3.4 Explain why carboxylic acids fizz with a carbonate but alcohols do not. [2]

Solutions

2.1 $2\text{CH}_3\text{COOH} + \text{Na}_2\text{CO}_3 \rightarrow 2\text{CH}_3\text{COONa} + \text{H}_2\text{O} + \text{CO}_2$.

2.2 add a carbonate (e.g. sodium carbonate); effervescence / fizzing of CO_2 is seen.

3.1 B. LiAlH_4 reduces a carboxylic acid to a primary alcohol.

3.2 B. A reactive metal gives a salt and hydrogen.

3.3 (a) $\text{CH}_3\text{COOH} + \text{NaOH} \rightarrow \text{CH}_3\text{COONa} + \text{H}_2\text{O}$; neutralisation.

(b) acidified $\text{K}_2\text{Cr}_2\text{O}_7$ (or KMnO_4); heat under reflux.

3.4 carboxylic acids are acidic — they release H^+ which reacts with the carbonate to give CO_2 ; alcohols are not acidic enough to react with a carbonate.