A-LEVEL CHEMISTRY • EXERCISE SHEET

13.4 Isomerism: structural isomerism and stereoisomerism

Name: _____ Class: _____ Date: _____ **Total: 12 marks**

KEY WORDS isomerism 异构 • structural 结构 • stereoisomerism 立体异构 • structural isomerism 结构异构 • chiral 手性

Objective

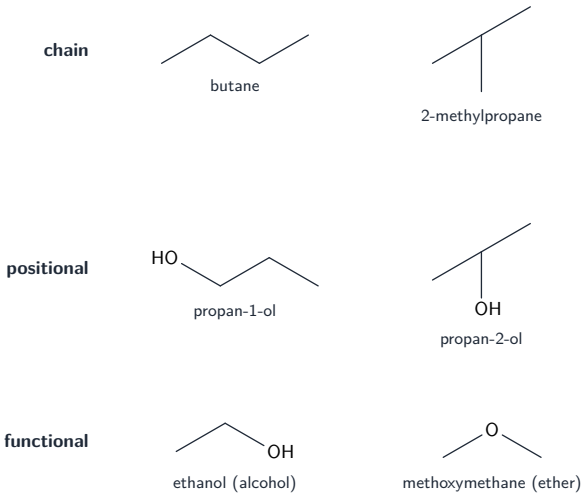
Build the skills to answer exam questions on **isomerism** 异构 — **structural** 结构 isomerism (chain, positional, functional group) and **stereoisomerism** 立体异构 (cis/trans and optical).

You must be able to:

- describe structural isomerism and its three kinds
- explain cis/trans isomerism in terms of restricted rotation at a C=C bond
- identify chiral centres and the geometrical isomerism in a given structure

1 Worked examples

■ Structural isomerism



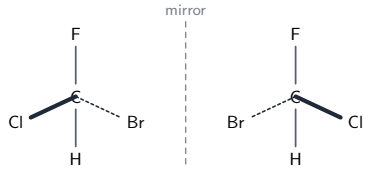
Chain, positional and functional-group isomerism — same molecular formula, different structure

■ Stereoisomerism

cis: both CH_3 on the same side

trans: CH_3 on opposite sides

Cis/trans isomerism at a C=C: the π bond stops rotation, fixing groups on the same (cis) or opposite (trans) side



enantiomers: mirror images that cannot be superimposed
(the central carbon has 4 different groups)

*A **chiral centre** 手性中心 (carbon with four different groups) gives two mirror-image **enantiomers** 对映体*

2 Practice

2.1 Name the three kinds of structural isomerism. [3]

2.2 State what is meant by a **chiral centre**. [1]

3 Exam-style questions

3.1 Cis/trans isomerism is possible because: [1]

- **A** single bonds cannot rotate
- **B** the π bond in C=C prevents rotation
- **C** the molecule is ionic
- **D** the carbons are sp^3 hybridised

3.2 Which compound has a chiral centre? [1]

- **A** $CH_3CH_2CH_3$
- **B** $CH_3CHBrCH_2CH_3$
- **C** CH_3CH_2OH
- **D** CH_2Cl_2

3.3 (a) Draw the cis and trans isomers of but-2-ene, $CH_3CH{=}CHCH_3$. [2]

(b) Explain why but-2-ene shows cis/trans isomerism but but-1-ene does not. [2]

3.4 2-hydroxypropanoic acid is $CH_3CH(OH)COOH$. State, with a reason, whether it has a chiral centre. [2]

Solutions

2.1 chain; positional; functional group isomerism.

2.2 a carbon atom bonded to four different groups.

3.1 B. The π bond in the C=C prevents rotation.

3.2 B. The second carbon in $CH_3CHBrCH_2CH_3$ has four different groups.

3.3 (a) cis — both CH_3 groups on the same side of the C=C; trans — on opposite sides. (b) but-2-ene has a different group (H and CH_3) on each double-bond carbon, so cis/trans is possible; but-1-ene has two H atoms on the end carbon, so the groups can be swapped without making a new isomer.

3.4 yes; the second carbon is bonded to four different groups (CH_3, OH, H, COOH), so it is a chiral centre.