Equilibria

A-Level Chemistry · Vocabulary in context

Word	Example
conjugate base /ˈkɒndʒuːgeɪt beɪs/	Cl− is the **conjugate base** of HCl. Cl− HCl
conjugate acid /ˈkɒndʒuːgeɪt ˈæsɪd/	NH4+ is the **conjugate acid** of NH3. NH4+ NH3
conjugate acid–base pair /ˈkɒndʒuːgeɪt ˈæsɪd beɪs peə/	HCl and Cl− are a **conjugate acid–base pair**. HCl Cl−
amphoteric /ˌæmfəʊˈterɪk/	Aluminium oxide is **amphoteric**: it reacts with acids and with bases.
acid dissociation constant /ˈæsɪd dɪˌsəʊsɪˈeɪʃn ˈkɒnstənt/	A weak acid has a small **acid dissociation constant**.
ionic product of water /aɪˈɒnɪk ˈprɒdʌkt ɒv ˈwɔːtə/	The **ionic product of water** is 1.0×10^{-14} mol^2 dm^{-6} at 25 °C. 25 °C 1.0×10^{-14} mol^2 dm^{-6}
strong acid /strɒŋ ˈæsɪd/	Hydrochloric acid is a **strong acid**.
strong alkali /strɒŋ ˈælkəlaɪ/	Sodium hydroxide is a **strong alkali**.
weak acid /wiːk ˈæsɪd/	Vinegar contains a **weak acid**.
buffer solution /ˈbʌfə səˈluːʃn/	A **buffer solution** keeps the pH almost the same. pH
solubility product /ˌsɒljuːˈbɪlɪti ˈprɒdʌkt/	Silver chloride has a very small **solubility product**.
common ion effect /ˈkɒmən ˈaɪɒn ɪˈfekt/	Adding salt makes less silver chloride dissolve: the **common ion effect**.

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partition coefficient /pɑːˈtɪʃn ˌkəʊɪˈfɪʃənt/	The **partition coefficient** compares how well iodine dissolves in two liquids.
solute /ˈsɒljuːt/	Salt is the **solute** in salt water.
polarity /pəʊˈlærɪti/	The **polarity** of water helps it dissolve salt.
solvent /ˈsɒlvənt/	Water is the **solvent** in salt water.

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