Electrochemistry

A-Level Chemistry · Vocabulary in context

Word	Example
electrolysis /ɪlekˈtrɒləsɪs/	**Electrolysis** of brine makes chlorine.
electrolyte /ɪˈlektrəlaɪt/	Salt water is an **electrolyte**: ions carry the current.
electrode /ɪˈlektrəʊd/	The positive **electrode** is the anode.
cathode /ˈkæθəʊd/	Electrons leave the hot **cathode**.
anode /ˈænəʊd/	The **anode** is the positive electrode.
Faraday constant /ˈfærədeɪ ˈkɒnstənt/	The **Faraday constant** is 96 500 C per mole of electrons. 96 500 C
Avogadro constant /ˌævəˈgædrəʊ ˈkɒnstənt/	The **Avogadro constant** is the number of particles in one mole.
electrode potential /ɪˈlektrəʊd pəˈtenʃl/	Zinc has a negative **electrode potential**.
standard electrode potential /ˈstændəd ɪˈlektrəʊd pəˈtenʃl/	The **standard electrode potential** of copper is +0.34 V. +0.34 V
standard hydrogen electrode /ˈstændəd ˈhaɪdrədʒɪn ɪˈlektrəʊd/	The **standard hydrogen electrode** has a potential of 0 V. 0 V
salt bridge /sɒlt brɪdʒ/	A **salt bridge** joins the two half-cells.
standard cell potential /ˈstændəd sel pəˈtenʃl/	The **standard cell potential** of a zinc–copper cell is 1.10 V. – 1.10 V

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oxidising agent /ˈɒksɪdaɪzɪŋ ˈeɪdʒənt/	Oxygen is an **oxidising agent**.
reducing agent /rɪˈdjuːsɪŋ ˈeɪdʒənt/	Carbon is a **reducing agent** in a blast furnace.
half-equation /hɑːf ɪˈkweɪʒn/	The **half-equation** $Cu^{2+} + 2e^- \rightarrow Cu$ shows reduction. $Cu^{2+} + 2e^- \rightarrow Cu$
feasibility /ˌfiːzɪˈbɪlɪti/	ΔG tells us about the **feasibility** of a reaction. ΔG
Nernst equation /nɜːnst ɪˈkweɪʒn/	The **Nernst equation** shows how the electrode potential changes with concentration.