Electrochemistry

A-Level Chemistry · Vocabulary in context

Word	Example
electrolysis /ɪlek'trɒləsɪs/	**Electrolysis** of brine makes chlorine.
electrolyte /ɪ'lektrəlaɪt/	Salt water is an **electrolyte**: ions carry the current.
electrode /ɪ'lektrəʊd/	The positive **electrode** is the anode.
cathode /'kæθəʊd/	Electrons leave the hot **cathode**.
anode /'ænəʊd/	The **anode** is the positive electrode.
Faraday constant /'færədeɪ 'kɒnstənt/	The **Faraday constant** is 96 500 C per mole of electrons. 96 500 C
Avogadro constant /ˌævə'gædrəʊ 'kɒnstənt/	The **Avogadro constant** is the number of particles in one mole.
electrode potential /ɪ'lektrəʊd pə'tenʃl/	Zinc has a negative **electrode potential**.
standard electrode potential /'stændəd ɪ'lektrəʊd pə'tenʃl/	The **standard electrode potential** of copper is +0.34 V. +0.34 V
standard hydrogen electrode /'stændəd 'haɪdrədʒn ɪ'lektrəʊd/	The **standard hydrogen electrode** has a potential of 0 V. 0 V
salt bridge /sɒlt brɪdʒ/	A **salt bridge** joins the two half-cells.
standard cell potential /'stændəd sel pə'tenʃl/	The **standard cell potential** of a zinc–copper cell is 1.10 V. – 1.10 V

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↑ **oxidising agent** /ˈɒksɪdaɪzɪŋ ˈeɪdʒənt/	Oxygen is an **oxidising agent**.
↓ **reducing agent** /rɪˈdjuːsɪŋ ˈeɪdʒənt/	Carbon is a **reducing agent** in a blast furnace.
÷ **half-equation** /hɑːf ɪˈkweɪʒn/	The **half-equation** Cu2+ + 2e− → Cu shows reduction. Cu2+ + 2e− → Cu
✓ **feasibility** /ˌfiːzɪˈbɪlɪti/	ΔG tells us about the **feasibility** of a reaction. ΔG
ƒ **Nernst equation** /nɜːnst ɪˈkweɪʒn/	The **Nernst equation** shows how the electrode potential changes with concentration.

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