Chemical energetics

A-Level Chemistry · Vocabulary in context

Word	Example
enthalpy change /enˈθælpi tʃeɪndʒ/	The **enthalpy change** of this reaction is −50 kJ/mol.
enthalpy change of atomisation /enˈθælpi tʃeɪndʒ ɒv əˌtɒmaɪˈzeɪʃn/	The **enthalpy change of atomisation** of sodium makes gas atoms from the solid.
lattice energy /ˈlætɪs ˈenədʒi/	Magnesium oxide has a large **lattice energy**.
electron affinity /ɪˈlektrɒn əˈfɪnɪti/	Chlorine has a negative first **electron affinity**: energy is released.
Born–Haber cycle /bɔːn ˈheɪbə ˈsaɪkl/	A **Born–Haber cycle** finds the lattice energy of sodium chloride.
enthalpy change of hydration /enˈθælpi tʃeɪndʒ ɒv haɪˈdreɪʃn/	Small ions have a large **enthalpy change of hydration**.
enthalpy change of solution /enˈθælpi tʃeɪndʒ ɒv səˈluːʃn/	Ammonium nitrate has a positive **enthalpy change of solution**: the water gets cold.
entropy /ˈentrəpi/	A file with high **entropy** is hard to compress.
entropy change /ˈentrəpi tʃeɪndʒ/	Melting ice has a positive **entropy change**: the particles become more disordered.
Gibbs free energy /gɪbz friː ˈenədʒi/	A reaction can happen if the **Gibbs free energy** change is negative.
feasibility /ˌfiːzɪˈbɪlɪti/	ΔG tells us about the **feasibility** of a reaction. 

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