The Periodic Table: chemical periodicity

A-Level Chemistry · Vocabulary in context

Word	Example
periodicity /ˌpɪərɪˌɒˈdɪsɪti/	**Periodicity** means patterns repeat across each period.
period /ˈpɪərɪəd/	The **period** of the wave is 0.02 s. 0.02 s
atomic radius /əˈtɒmɪk ˈreɪdɪəs/	Sodium has a larger **atomic radius** than chlorine.
ionic radius /aɪˈɒnɪk ˈreɪdɪəs/	The Na+ ion has a smaller **ionic radius** than the sodium atom. Na+
melting point /ˈmeltɪŋ pɔɪnt/	The **melting point** of ice is 0 °C. 0 °C
electrical conductivity /ɪˈlektrɪkl kɒndəkˈtɪvɪti/	Copper has a high **electrical conductivity**.
giant metallic /ˈdʒaɪənt məˈtælɪk/	Copper has a **giant metallic** structure.
giant molecular /ˈdʒaɪənt məˈlekjʊlə/	Diamond has a **giant molecular** structure.
simple molecular /ˈsɪmpl məˈlekjʊlə/	Iodine has a **simple molecular** structure.
van der Waals' forces /ˈvændɜː wɑːlz ˈfɔːsɪz/	**Van der Waals' forces** hold iodine molecules together.
steam /stiːm/	Ethene reacts with **steam** to make ethanol.
oxidation number /ˌɒksɪˈdeɪʃn ˈnʌmbə/	In H2O, oxygen has an **oxidation number** of −2. H2O −2
valence shell /ˈveɪləns ʃel/	Sodium has one electron in its **valence shell**.

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electron /ɪˈlektrɒn/	An **electron** has a negative charge.
oxide /ˈɒksaɪd/	Magnesium **oxide** forms when magnesium burns.
amphoteric /ˌæmfəʊˈterɪk/	Aluminium oxide is **amphoteric**: it reacts with acids and with bases.
hydroxide /haɪˈdrɒksaɪd/	Sodium **hydroxide** is a strong alkali.
polarises /ˈpəʊləraɪzɪz/	The small Al3+ ion **polarises** the water molecules around it. Al3+
hydrolysis /haɪˈdrɒləsɪs/	**Hydrolysis** uses water to break a large molecule into smaller ones.
electronegativity /ɪˌlektrəʊŋgəˈtɪvɪti/	Fluorine has the highest **electronegativity** of all elements.

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