Electrochemistry

A-Level Chemistry · Vocabulary in context

Word	Example
# **oxidation number** /ˌɒksɪˈdeɪʃn ˈnʌmbə/	In H2O, oxygen has an **oxidation number** of −2. H2O −2
↑ **oxidation** /ˌɒksɪˈdeɪʃn/	**Oxidation** is the loss of electrons.
electron /ɪˈlektrɒn/	An **electron** has a negative charge.
↑↓ **redox** /rɪˈdɒks/	A **redox** reaction has oxidation and reduction together.
↓ **reduction** /rɪˈdʌkʃn/	**Reduction** is the gain of electrons.
⇄ **electron transfer** /ɪˈlektrɒn ˈtrænsfɜː/	**Electron transfer** happens when magnesium reacts with copper ions.
↖↗ **half-equations** /hɑːf ɪˈkweɪʒnz/	Write two **half-equations**: one for oxidation and one for reduction.
↖↗ **disproportionation** /ˌdɪsprəˈpɔːʃəneɪʃn/	In **disproportionation**, the same element is oxidised and reduced.
↑ **oxidising agent** /ˈɒksɪdaɪzɪŋ ˈeɪdʒənt/	Oxygen is an **oxidising agent**.
↓ **reducing agent** /rɪˈdjuːsɪŋ ˈeɪdʒənt/	Carbon is a **reducing agent** in a blast furnace.

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