

23 Lattice energy and electron affinity – Part 1

Name: _____ Class: _____ Date: _____

Vocabulary

Write each word three times.

English	Write it 3 times (English)
enthalpy changes	_____
lattice energy	_____
enthalpy change of atomisation	_____
electron affinity	_____

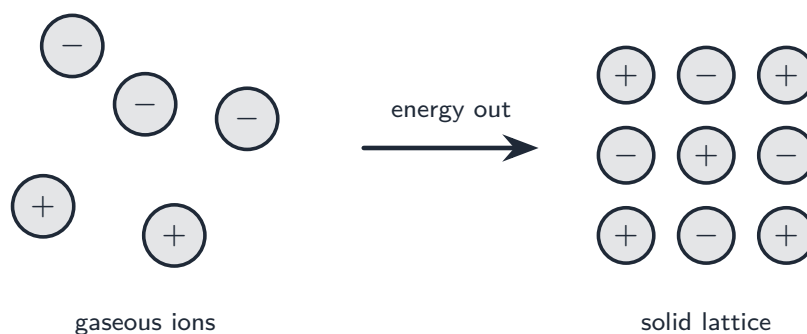
Recall

Born–Haber cycles use **enthalpy changes** (ΔH). Two new ones describe forming gaseous atoms and the ionic solid.

New ideas

Read these first, then do the Practice.

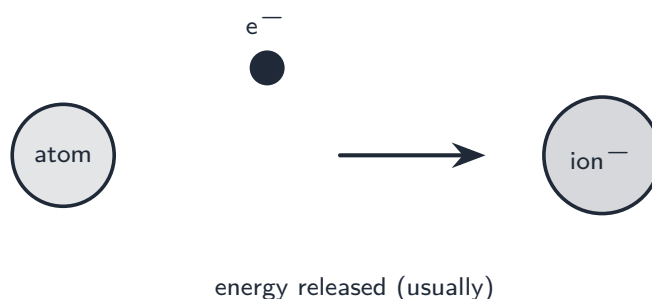
■ 1 Lattice energy



The **lattice energy** is the energy change when one mole of an ionic solid forms from its gaseous ions. It is always *negative* (strong bonds form). The **enthalpy change of atomisation** (to make gaseous atoms) is always positive.

Worked example. For sodium chloride, gaseous Na^+ and Cl^- ions come together into the solid lattice and release energy, so its lattice energy is a large *negative* value (about -787 kJ mol^{-1}).

■ 2 Electron affinity



The first **electron affinity** is the energy change when gaseous atoms each gain one electron to form 1^- ions. It is usually negative.

Watch out: lattice energy is *negative* (ions attract, energy out); atomisation and ionisation are *positive* (energy in). The *first* electron affinity is usually negative, but the *second* is positive – you must push an electron onto an already-negative ion.

Practice

Try these in order.

23.1 Is the *lattice energy* positive or negative? [1]

23.2 Why does that sign make sense? [1]

23.3 What is the *enthalpy change of atomisation*? [1]

23.4 What is the first *electron affinity*? [1]

23.5 Is the first electron affinity usually positive or negative? [1]

23.6 Challenge. Magnesium oxide ($\text{Mg}^{2+}\text{O}^{2-}$) has a much more negative lattice energy than sodium chloride (Na^+Cl^-). Explain why, in terms of ionic charge. [2]
