

9 Periodicity (Period 3) – Part 2

Name: _____ Class: _____ Date: _____

Vocabulary

Write each word three times.

English	Write it 3 times (English)
oxides	_____
chlorides	_____
basic	_____
amphoteric	_____
acidic	_____
hydrolyse	_____

Recall

From Part 1 you know Period 3 goes from metals (left) to non-metals (right). Here we see how that changes the **oxides** and **chlorides**.

- Metal oxides are **basic**; non-metal oxides are **acidic**.
- One oxide in the middle is special: it reacts both ways.

New ideas

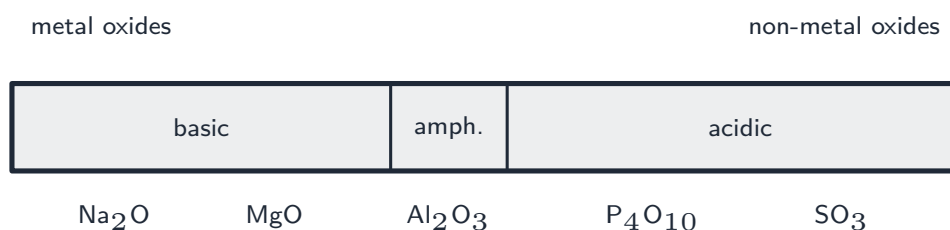
Read these first, then do the Practice.

■ 1 Acid–base nature of the oxides

Across Period 3 the oxides change from **basic** to **acidic**:

- **metal oxides** (Na_2O , MgO) are **basic**.
- **Al_2O_3 is amphoteric** – it reacts with **both** acids and bases.

- **non-metal oxides** (P_4O_{10} , SO_3) are **acidic** .



■ 2 Aluminium oxide is amphoteric



It behaves as a base with acids, and as an acid with bases – that is what **amphoteric** means.

■ 3 Chlorides with water

- **ionic chlorides** (NaCl , MgCl_2) simply **dissolve**, giving a near-neutral solution.
- **covalent chlorides** (SiCl_4 , PCl_5) **hydrolyse** (react with water), giving an **acidic** solution and **HCl** fumes.

Watch out: the oxides change from basic (metals) through amphoteric (Al_2O_3) to acidic (non-metals) – the trend follows the metal-to-non-metal change across the period.

Practice

Try these in order.

9.1 State whether each oxide is basic or acidic: (a) Na_2O , (b) SO_3 . [2]

9.2 What does *amphoteric* mean? [1]

9.3 Which Period 3 oxide is amphoteric?

[1]

9.4 Sodium chloride dissolves in water to a neutral solution, but phosphorus(V) chloride reacts with water. Explain this difference.

[2]

9.5 Name the gas given off (as fumes) when a covalent chloride is added to water.

[1]

9.6 Challenge. Predict which would raise the pH of water – magnesium oxide (MgO) or sulfur trioxide (SO₃) – and explain.

[2]