

22.2 Mass spectrometry

Name: _____ Class: _____ Date: _____

Total: 9 marks**KEY WORDS** molecular ion

• relative atomic mass

• molecular ion peak

• isotopic abundance

• mass spectrometry

Objective

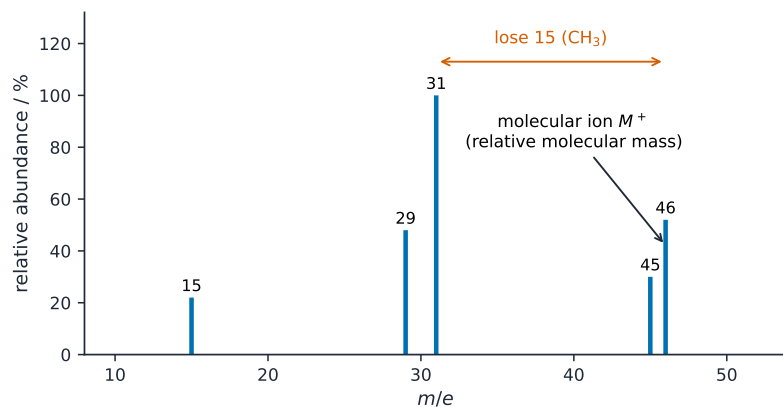
Build the skills to answer exam questions on **mass spectrometry** – the **molecular ion**, **fragmentation**, the **[M+1]** carbon count, and the **[M+2]** halogen test.

You must be able to:

- calculate relative atomic mass from isotopic abundances or a mass spectrum
- deduce molecular mass from the molecular ion peak and identify fragments
- use the [M+1] peak to find the number of carbons, and [M+2] for Cl/Br

1 Key knowledge and worked examples

■ Relative atomic mass and the molecular ion

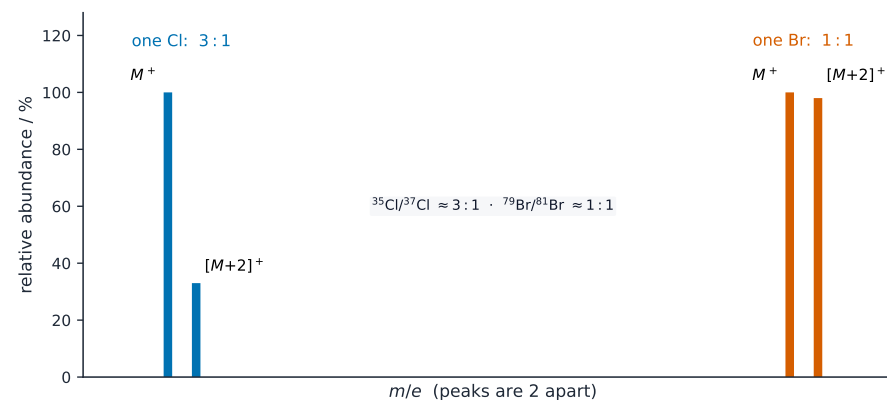


M^+ peak $\approx M_r$ · gaps between peaks = mass of fragments lost

The peaks are at different m/e values; the **molecular ion** M^+ (highest m/e) gives the M_r

$$A_r = \frac{\sum(\text{isotope mass} \times \text{abundance})}{\sum \text{abundance}}$$

■ Fragmentation, [M+1] and [M+2]



A small [M+1] peak comes from ^{13}C ; an [M+2] peak shows a Cl or Br atom

- fragments:** the molecule breaks into smaller ions; the mass lost identifies the group (e.g. loss of 15 = CH_3).
- [M+1] carbon count:** $n = \frac{100 \times \text{abundance of } [M+1]}{1.1 \times \text{abundance of } M^+}$.
- [M+2]:** a peak two higher $\rightarrow$ Cl (3:1) or Br (1:1).

2 Practice

2.1 State what the molecular ion peak (highest m/e) tells you about a compound. [1]

2.2 A mass spectrum shows a peak at M and another of nearly equal height at $M+2$. What element is present? [1]

3 Exam-style questions

3.1 Chlorine has isotopes ^{35}Cl (75%) and ^{37}Cl (25%). Its relative atomic mass is: [1]

☐

A 35.0

B 35.5

C 36.0

D 37.0

3.2 Loss of a fragment of mass 15 from a molecular ion suggests the loss of: [1]

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A OH

B CH_3

C C_2H_5

D Cl

3.3 A compound has a molecular ion peak M^+ at $m/e = 88$ (abundance 100) and an $[M+1]$ peak (abundance 4.4).

(a) Use $n = \frac{100 \times \text{abundance of } [M+1]}{1.1 \times \text{abundance of } M^+}$ to find the number of carbon atoms. [2]

(b) The compound is a carboxylic acid of M_r 88. Suggest its molecular formula. [1]

3.4 A mass spectrum has peaks at $m/e = 78$ and 80 in a 1:1 ratio. State, with a reason, which halogen is present. [2]
