

. Answer the following questions about the compounds NH_2Cl and NCl_3 . The Lewis electron-dot diagrams of the two compounds are shown.



(a) Calculate the number of moles of NH_2Cl (molar mass 51.48 g/mol) present in 1.0 L of a solution in which the concentration of NH_2Cl is 0.0016 g/L.

(b) NH_2Cl is highly soluble in water, whereas NCl_3 is nearly insoluble. Explain this observation in terms of the types and relative strengths of the intermolecular forces between each of the solutes and water.

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Continue your response to **QUESTION 4** on this page.

(c) The value of $\Delta H^\circ_{\text{vaporization}}$ for $\text{NCl}_3(l)$ is 32.9 kJ/mol. Calculate the amount of energy required to vaporize a 15.0 g sample of NCl_3 (molar mass 120.36 g/mol).

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