

- 3 (a) (i) To a 1 cm depth of **FB 4** in a boiling tube, slowly add a 1 cm depth of dilute nitric acid. Stir gently until the reaction is complete. Use this solution for your tests. Select reagents to identify which sodium halide is present. Record details of the reagents used and your observations. Identify the sodium halide that is present as an impurity in **FB 4**.

The formula of the impurity is .....

[2]

- (ii) Suggest why it is necessary to add the nitric acid to **FB 4** before carrying out your tests in (a)(i).

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..... [1]

- (b) **FB 7** contains a metal cation which is listed in the Qualitative analysis notes.

- (i) Place half the sample of **FB 7** in a hard-glass test-tube. Heat gently at first and then more strongly. Record your observations.

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..... [2]

- (ii) Use the 25 cm<sup>3</sup> measuring cylinder to measure 20 cm<sup>3</sup> of **FB 1**. Place the remaining sample of **FB 7** in a boiling tube and add portions of **FB 1** from the measuring cylinder until all the **FB 7** has reacted. Stir until the reaction is complete. Record your observations.

**Keep this solution for use in (b)(iii).**

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..... [2]

- (iii) Carry out the following tests. For each test use a 1 cm depth of the solution from (b)(ii) in a test-tube.

Record your observations in Table 3.1.

Identify the metal ion present in **FB 7**.

**Table 3.1**

<i>test</i>	<i>observations</i>
<b>Test 1</b> Add aqueous sodium hydroxide.	
<b>Test 2</b> Add aqueous ammonia.	
<b>Test 3</b> Add a piece of aluminium foil.	

The formula of the metal ion in **FB 7** is .....

[4]

- (c) **FB 8** is an acidified aqueous solution containing a metal cation which is listed in the Qualitative analysis notes.

- (i) Carry out the following tests. For each test use a 1 cm depth of **FB 8** in a test-tube. Record your observations in Table 3.2. Identify the metal ion present in **FB 8**.

**Table 3.2**

<i>test</i>	<i>observations</i>
<b>Test 1</b> Add aqueous sodium hydroxide.	
<b>Test 2</b> Add a 1 cm depth of acidified aqueous potassium manganate(VII).	

The formula of the metal ion in **FB 8** is .....

[3]

- (ii) State the type of reaction between the metal ion in **FB 8** and acidified aqueous potassium manganate(VII). Explain how your observations support your answer.

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..... [1]

[Total: 15]

### Qualitative analysis notes

#### 1 Reactions of cations

cation	reaction with	
	NaOH(aq)	NH <sub>3</sub> (aq)
aluminium, Al <sup>3+</sup> (aq)	white ppt. soluble in excess	white ppt. insoluble in excess
ammonium, NH <sub>4</sub> <sup>+</sup> (aq)	no ppt. ammonia produced on warming	–
barium, Ba <sup>2+</sup> (aq)	faint white ppt. is observed unless [Ba <sup>2+</sup> (aq)] is very low	no ppt.
calcium, Ca <sup>2+</sup> (aq)	white ppt. unless [Ca <sup>2+</sup> (aq)] is very low	no ppt.
chromium(III), Cr <sup>3+</sup> (aq)	grey-green ppt. soluble in excess giving dark green solution	grey-green ppt. insoluble in excess
copper(II), Cu <sup>2+</sup> (aq)	pale blue ppt. insoluble in excess	pale blue ppt. soluble in excess giving dark blue solution
iron(II), Fe <sup>2+</sup> (aq)	green ppt. turning brown on contact with air insoluble in excess	green ppt. turning brown on contact with air insoluble in excess
iron(III), Fe <sup>3+</sup> (aq)	red-brown ppt. insoluble in excess	red-brown ppt. insoluble in excess
magnesium, Mg <sup>2+</sup> (aq)	white ppt. insoluble in excess	white ppt. insoluble in excess
manganese(II), Mn <sup>2+</sup> (aq)	off-white ppt. rapidly turning brown on contact with air insoluble in excess	off-white ppt. rapidly turning brown on contact with air insoluble in excess
zinc, Zn <sup>2+</sup> (aq)	white ppt. soluble in excess	white ppt. soluble in excess

#### 2 Reactions of anions

anion	reaction
carbonate, CO <sub>3</sub> <sup>2-</sup>	CO <sub>2</sub> liberated by dilute acids
chloride, Cl <sup>-</sup> (aq)	gives white ppt. with Ag <sup>+</sup> (aq) (soluble in NH <sub>3</sub> (aq))
bromide, Br <sup>-</sup> (aq)	gives cream / off-white ppt. with Ag <sup>+</sup> (aq) (partially soluble in NH <sub>3</sub> (aq))
iodide, I <sup>-</sup> (aq)	gives pale yellow ppt. with Ag <sup>+</sup> (aq) (insoluble in NH <sub>3</sub> (aq))
nitrate, NO <sub>3</sub> <sup>-</sup> (aq)	NH <sub>3</sub> liberated on heating with OH <sup>-</sup> (aq) and Al foil
nitrite, NO <sub>2</sub> <sup>-</sup> (aq)	NH <sub>3</sub> liberated on heating with OH <sup>-</sup> (aq) and Al foil; decolourises acidified aqueous KMnO <sub>4</sub>
sulfate, SO <sub>4</sub> <sup>2-</sup> (aq)	gives white ppt. with Ba <sup>2+</sup> (aq) (insoluble in excess dilute strong acids); gives white ppt. with high [Ca <sup>2+</sup> (aq)]
sulfite, SO <sub>3</sub> <sup>2-</sup> (aq)	gives white ppt. with Ba <sup>2+</sup> (aq) (soluble in excess dilute strong acids); decolourises acidified aqueous KMnO <sub>4</sub>
thiosulfate, S <sub>2</sub> O <sub>3</sub> <sup>2-</sup> (aq)	gives off-white / pale yellow ppt. slowly with H <sup>+</sup>

#### 3 Tests for gases

gas	test and test result
ammonia, NH <sub>3</sub>	turns damp red litmus paper blue
carbon dioxide, CO <sub>2</sub>	gives a white ppt. with limewater
hydrogen, H <sub>2</sub>	'pops' with a lighted splint
oxygen, O <sub>2</sub>	relights a glowing splint