

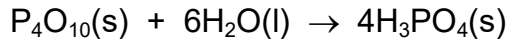
9 The data shown are needed for this question.

$$\Delta H_f^\ominus(\text{P}_4\text{O}_{10}(\text{s})) = -3012 \text{ kJ mol}^{-1}$$

$$\Delta H_f^\ominus(\text{H}_2\text{O}(\text{l})) = -286 \text{ kJ mol}^{-1}$$

$$\Delta H_f^\ominus(\text{H}_3\text{PO}_4(\text{s})) = -1279 \text{ kJ mol}^{-1}$$

What is ΔH^\ominus for the reaction shown?



A $-9844 \text{ kJ mol}^{-1}$

B -388 kJ mol^{-1}

C -97 kJ mol^{-1}

D $+2019 \text{ kJ mol}^{-1}$