

- 9** When 0.47 g of a hydrocarbon is completely burnt in air, the energy released heats 200 g of water from 23.7 °C to 41.0 °C.

What is the amount of energy absorbed, in Joules, by the water?

- A** $0.47 \times 4.18 \times 17.3$
- B** $0.47 \times 4.18 \times (273 + 17.3)$
- C** $200 \times 4.18 \times 17.3$
- D** $200 \times 4.18 \times (273 + 17.3)$